Name
Date **Period**

Grade:

Lab 17
CONSERVATION OF MATTER

PRELAB QUESTIONS

1. Explain the Law of Conservation of Matter.
2. Define: reactants, products, chemical equation.
3. Write a balanced chemical equation for the reaction of aqueous sodium carbonate and aqueous calcium chloride. Label the reactants and products. Explain how this illustrates the conservation of matter in regards to the atoms in the equation.

NAME _____ PERIOD_____
DATE _____ LAB PARTNERS

EXPERIMENT 17
CONSERVATION OF MATTER

QUALITATIVE OBSERVATIONS

Step 5:

Step 7:

QUANTITATIVE DATA

Mass A (before mixing)	g
Mass B (after mixing Na_2CO_3 and $CaCl_2$)	g
Mass C (after mixing H_2SO_4 with contents from step 5)	g

CONCLUSION QUESTIONS

1. In procedure 5, what observations occurred that shows a chemical reaction took place? In procedure 7?

2. Why do you leave the flask unstoppered after H_2SO_4 solution was added?

3. Account for any differences in the masses of A, B, and C

4. Discuss how the Law of Conservation of Matter relates to balancing chemical equations.

5. Does this exercise verify the Law of Conservation of matter? Explain

Discussion

Conclusion