

Reactions 3

AP Chemistry

This exercise must be completed and brought to my room before the start of first period on Friday. Failure to do so will incur a 25% penalty unless there is a legal reason.

The Reactions are worth 3 points apiece. 1 point for the correct reactants and 2 points for the correct products. The maximum score you can achieve is 24/20.

Instructions

Give the formulas to show the reactants and the products of the following chemical reactions. Each of the reactions occurs in aqueous solution unless otherwise indicated. Represent substances in solution as ions if the substance is extensively ionized. Omit formulas for any ions or molecules that are unchanged by the reaction (net ionic). In all cases a reaction occurs.

Section 1

- a) Solid sodium hydrogen carbonate (sodium bicarbonate) is strongly heated
- b) Carbon disulfide vapor is burned in excess oxygen.
- c) Solid cesium oxide is added to water.
- d) Solutions of zinc sulfate and sodium phosphate are mixed.
- e) A solution of tin(II)chloride is added to an acidified solution of potassium permanganate.
- f) A solution of ammonium thiocyanate is added to a solution of iron(III)chloride.
- g) Samples of boron trichloride gas and ammonia gas are mixed.
- h) A stream of chlorine gas is passed through a solution of cold, dilute sodium hydroxide.