

Reactions 2

AP Chemistry

This exercise must be completed and brought to my room before the start of first period on Friday. Failure to do so will incur a 25% penalty unless there is a legal reason.

The Reactions are worth 3 points apiece. 1 point for the correct reactants and 2 points for the correct products. The maximum score you can achieve is 24/20.

Instructions

Give the formulas to show the reactants and the products of the following chemical reactions. Each of the reactions occurs in aqueous solution unless otherwise indicated. Represent substances in solution as ions if the substance is extensively ionized. Omit formulas for any ions or molecules that are unchanged by the reaction (net ionic). In all cases a reaction occurs.

- (a) A solution of ammonium thiocyanate is added to a solution of iron (III) chloride
- (b) Solid calcium is added to warm water.
- (c) Hydrogen sulfide gas is bubbled through a solution of potassium hydroxide.
- (d) Hydrogen iodide gas is bubbled into a solution of lithium carbonate.
- (e) An acidic solution of potassium dichromate is added to a solution of iron (II) nitrate.
- (f) Excess potassium hydroxide solution is added to a solution of potassium dihydrogen phosphate.
- (g) A solution of sodium phosphate is added to a solution of aluminum nitrate.
- (h) Propanol is burned completely in air.